

# Ch. 1 - Matter

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## I. States of Matter (p.12)

- ♦ Kinetic Molecular Theory
- ♦ States of Matter

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## A. Kinetic Molecular Theory

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### ♦ KMT

- ♦ Particles of matter are always in motion.
- ♦ The kinetic energy (speed) of these particles increases as temperature increases.


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## B. Four States of Matter

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### ♦ Solids

- ♦ very low KE - particles vibrate but can't move around
- ♦ fixed shape
- ♦ fixed volume




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## B. Four States of Matter

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### ♦ Liquids

- ♦ low KE - particles can move around but are still close together
- ♦ variable shape
- ♦ fixed volume



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## B. Four States of Matter

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### ✦ Gases

- ♦ high KE - particles can separate and move throughout container
- ♦ variable shape
- ♦ variable volume



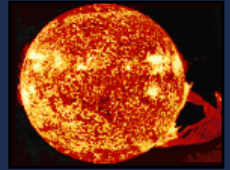
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## B. Four States of Matter

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### ✦ Plasma

- ♦ very high KE - particles collide with enough energy to break into charged particles (+/-)
- ♦ gas-like, variable shape & volume
- ♦ stars, fluorescent light bulbs, CRTs



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